

A grayscale photograph of a hand holding a pen, writing on lined paper. The text on the paper is mirrored, appearing as if written on a transparent surface. The words visible include 'things', 'giving', 'while', 'all', 'started', 'for', 'they', and 'the'.

Part II Practice Test 1

AP[®] Chemistry Exam

SECTION I: Multiple-Choice Questions

DO NOT OPEN THIS BOOKLET UNTIL YOU ARE TOLD TO DO SO.**At a Glance****Total Time**

1 hour and 30 minutes

Number of Questions

60

Percent of Total Grade

50%

Writing Instrument

Pencil required

Instructions

Section I of this examination contains 60 multiple-choice questions. Fill in only the ovals for numbers 1 through 60 on your answer sheet.

CALCULATORS MAY NOT BE USED IN THIS PART OF THE EXAMINATION.

Indicate all of your answers to the multiple-choice questions on the answer sheet. No credit will be given for anything written in this exam booklet, but you may use the booklet for notes or scratch work. After you have decided which of the suggested answers is best, completely fill in the corresponding oval on the answer sheet. Give only one answer to each question. If you change an answer, be sure that the previous mark is erased completely. Here is a sample question and answer.

Sample QuestionSample Answer

Chicago is a

(A) state

(B) city

(C) country

(D) continent

(A) ☒ (C) (D)

Use your time effectively, working as quickly as you can without losing accuracy. Do not spend too much time on any one question. Go on to other questions and come back to the ones you have not answered if you have time. It is not expected that everyone will know the answers to all the multiple-choice questions.

About Guessing

Many candidates wonder whether or not to guess the answers to questions about which they are not certain. Multiple-choice scores are based on the number of questions answered correctly. Points are not deducted for incorrect answers, and no points are awarded for unanswered questions. Because points are not deducted for incorrect answers, you are encouraged to answer all multiple-choice questions. On any questions you do not know the answer to, you should eliminate as many choices as you can, and then select the best answer among the remaining choices.

Disclaimer

This test is an approximation of the test that you will take. For up-to-date information, please remember to check the AP Students website.

CHEMISTRY

SECTION I

Time—1 hour and 30 minutes

INFORMATION IN THE TABLE BELOW AND ON THE FOLLOWING PAGES MAY BE USEFUL IN ANSWERING THE QUESTIONS IN THIS SECTION OF THE EXAMINATION.

PERIODIC TABLE OF THE ELEMENTS

DO NOT DETACH FROM BOOK.

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***Lanthanide Series**

† Actinide Series

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ADVANCED PLACEMENT CHEMISTRY EQUATIONS AND CONSTANTS

Throughout the test the following symbols have the definitions specified unless otherwise noted.

L, mL = liter(s), milliliter(s)
 g = gram(s)
 nm = nanometer(s)
 atm = atmosphere(s)

mm Hg = millimeters of mercury
 J, kJ = joule(s), kilojoule(s)
 V = volt(s)
 mol = mole(s)

ATOMIC STRUCTURE

$$E = h\nu$$

$$c = \lambda\nu$$

E = energy
 ν = frequency
 λ = wavelength

Planck's constant, $h = 6.626 \times 10^{-34}$ J sSpeed of light, $c = 2.998 \times 10^8$ m s⁻¹Avogadro's number = 6.022×10^{23} mol⁻¹Electron charge, $e = -1.602 \times 10^{-19}$ coulombs

EQUILIBRIUM

$$K_c = \frac{[C]^c[D]^d}{[A]^a[B]^b}, \text{ where } aA + bB \rightleftharpoons cC + dD$$

$$K_p = \frac{(P_C)^c(P_D)^d}{(P_A)^a(P_B)^b}$$

$$K_a = \frac{[H^+][A^-]}{[HA]}$$

$$K_b = \frac{[OH^-][HB^+]}{[B]}$$

$$K_w = [H^+][OH^-] = 1.0 \times 10^{-14} \text{ at } 25^\circ\text{C}$$

$$= K_a \times K_b$$

$$\text{pH} = -\log[H^+], \text{ pOH} = -\log[OH^-]$$

$$14 = \text{pH} + \text{pOH}$$

$$\text{pH} = \text{p}K_a + \log \frac{[A^-]}{[HA]}$$

$$\text{p}K_a = -\log K_a, \text{ p}K_b = -\log K_b$$

Equilibrium Constants

 K_c (molar concentrations) K_p (gas pressures) K_a (weak acid) K_b (weak base) K_w (water)

KINETICS

$$\ln[A]_t - \ln[A]_0 = -kt$$

$$\frac{1}{[A]_t} - \frac{1}{[A]_0} = kt$$

$$t_{1/2} = \frac{0.693}{k}$$

 k = rate constant t = time $t_{1/2}$ = half-life

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GASES, LIQUIDS, AND SOLUTIONS

$$PV = nRT$$

$$P_A = P_{\text{total}} \times X_A, \text{ where } X_A = \frac{\text{moles A}}{\text{total moles}}$$

$$P_{\text{total}} = P_A + P_B + P_C + \dots$$

$$n = \frac{m}{M}$$

$$K = ^\circ\text{C} + 273$$

$$D = \frac{m}{V}$$

$$KE \text{ per molecule} = \frac{1}{2}mv^2$$

Molarity, M = moles of solute per liter of solution

$$A = abc$$

P = pressure

V = volume

T = temperature

n = number of moles

m = mass

M = molar mass

D = density

KE = kinetic energy

v = velocity

A = absorbance

a = molar absorptivity

b = path length

c = concentration

Gas constant, $R = 8.314 \text{ J mol}^{-1} \text{ K}^{-1}$

$$= 0.08206 \text{ L atm mol}^{-1} \text{ K}^{-1}$$

$$= 62.36 \text{ L torr mol}^{-1} \text{ K}^{-1}$$

$$1 \text{ atm} = 760 \text{ mm Hg}$$

$$= 760 \text{ torr}$$

$$\text{STP} = 0.00^\circ\text{C and } 1.000 \text{ atm}$$

THERMOCHEMISTRY/ ELECTROCHEMISTRY

$$q = mc\Delta T$$

$$\Delta S^\circ = \sum S^\circ \text{ products} - \sum S^\circ \text{ reactants}$$

$$\Delta H^\circ = \sum \Delta H_f^\circ \text{ products} - \sum \Delta H_f^\circ \text{ reactants}$$

$$\Delta G^\circ = \sum \Delta G_f^\circ \text{ products} - \sum \Delta G_f^\circ \text{ reactants}$$

$$\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$$

$$= -RT \ln K$$

$$= -nFE^\circ$$

$$I = \frac{q}{t}$$

q = heat

m = mass

c = specific heat capacity

T = temperature

S° = standard entropy

H° = standard enthalpy

G° = standard free energy

n = number of moles

E° = standard reduction potential

I = current (amperes)

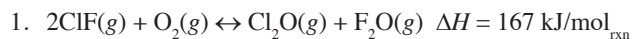
q = charge (coulombs)

t = time (seconds)

Faraday's constant, $F = 96,485$ coulombs per mole of electrons

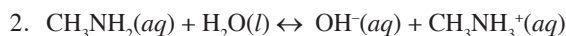
$$1 \text{ volt} = \frac{1 \text{ joule}}{1 \text{ coulomb}}$$

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During the reaction above, the product yield can be increased by increasing the temperature of the reaction. Why is this effective?

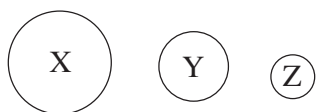
- (A) The reaction is endothermic; therefore adding heat will shift it to the right.
- (B) Increasing the temperature increases the speed of the molecules, meaning there will be more collisions between them.
- (C) The reactants are less massive than the products, and an increase in temperature will cause their kinetic energy to increase more than that of the products.
- (D) The increase in temperature allows for a higher percentage of molecular collisions to occur with the proper orientation to create the product.



The above equation represents the reaction between the base methylamine ($K_b = 4.38 \times 10^{-4}$) and water. Which of the following best represents the concentrations of the various species at equilibrium?

- (A) $[\text{OH}^-] > [\text{CH}_3\text{NH}_2] = [\text{CH}_3\text{NH}_3^+]$
- (B) $[\text{OH}^-] = [\text{CH}_3\text{NH}_2] = [\text{CH}_3\text{NH}_3^+]$
- (C) $[\text{CH}_3\text{NH}_2] > [\text{OH}^-] > [\text{CH}_3\text{NH}_3^+]$
- (D) $[\text{CH}_3\text{NH}_2] > [\text{OH}^-] = [\text{CH}_3\text{NH}_3^+]$

3. The diagram below shows the relative atomic sizes of three different elements from the same period. Which of the following statements must be true?



- (A) The effective nuclear charge will be the greatest in element X.
- (B) The first ionization energy will be greatest in element X.
- (C) The electron shielding effect will be greatest in element Z.
- (D) The electronegativity value will be greatest in element Z.

4. A sealed, rigid container contains three gases: 28.0 g of nitrogen, 40.0 g of argon, and 36.0 g of water vapor. If the total pressure exerted by the gases is 2.0 atm, what is the partial pressure of the nitrogen?

- (A) 0.33 atm
- (B) 0.40 atm
- (C) 0.50 atm
- (D) 2.0 atm

5. A sample of liquid NH_3 is brought to its boiling point. Which of the following occurs during the boiling process?

- (A) The N–H bonds within the NH_3 molecules break apart.
- (B) The overall temperature of the solution rises as the NH_3 molecules speed up.
- (C) The amount of energy within the system remains constant.
- (D) The hydrogen bonds holding separate NH_3 molecules together break apart.

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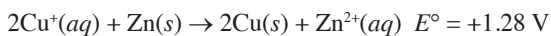
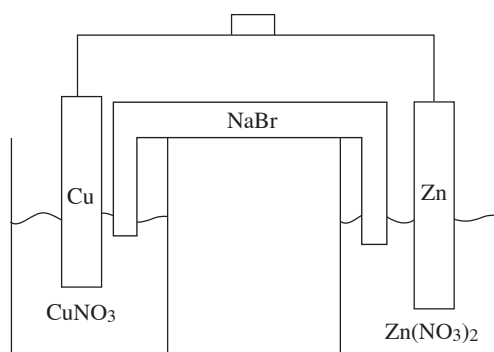
Questions 6-10 refer to the following.

Two half-cells are set up as follows:

Half-Cell A: Strip of $\text{Cu}(s)$ in $\text{CuNO}_3(aq)$

Half-Cell B: Strip of $\text{Zn}(s)$ in $\text{Zn}(\text{NO}_3)_2(aq)$

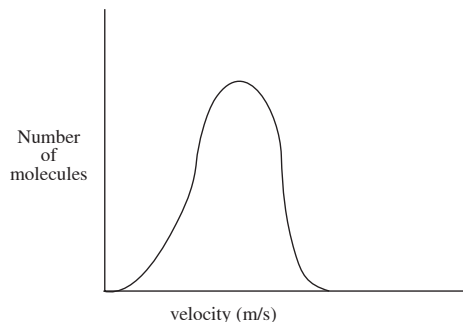
When the cells are connected according to the diagram below, the following reaction occurs:



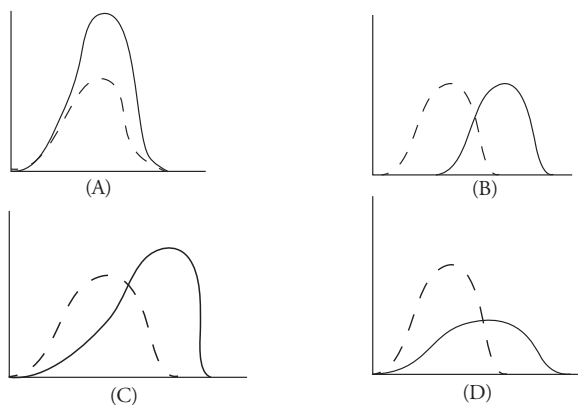
6. Correctly identify the anode and cathode in this reaction as well as where oxidation and reduction are taking place.
 - (A) Cu is the anode where oxidation occurs, and Zn is the cathode where reduction occurs.
 - (B) Cu is the anode where reduction occurs, and Zn is the cathode where oxidation occurs.
 - (C) Zn is the anode where oxidation occurs, and Cu is the cathode where reduction occurs.
 - (D) Zn is the anode where reduction occurs, and Cu is the cathode where oxidation occurs.
7. How many moles of electrons must be transferred to create 127 g of copper?
 - (A) 1 mole of electrons
 - (B) 2 moles of electrons
 - (C) 3 moles of electrons
 - (D) 4 moles of electrons
8. If the $\text{Cu}^+ + e^- \rightarrow \text{Cu}(s)$ half-reaction has a standard reduction potential of +0.52 V, what is the standard reduction potential for the $\text{Zn}^{2+} + 2e^- \rightarrow \text{Zn}(s)$ half-reaction?
 - (A) +0.76 V
 - (B) -0.76 V
 - (C) +0.24 V
 - (D) -0.24 V
9. As the reaction progresses, what will happen to the overall voltage of the cell?
 - (A) It will increase as $[\text{Zn}^{2+}]$ increases.
 - (B) It will increase as $[\text{Cu}^+]$ increases.
 - (C) It will decrease as $[\text{Zn}^{2+}]$ increases.
 - (D) The voltage will remain constant.
10. What will happen in the salt bridge as the reaction progresses?
 - (A) The Na^+ ions will flow to the Cu/Cu^+ half-cell.
 - (B) The Br^- ions will flow to the Cu/Cu^+ half-cell.
 - (C) Electrons will transfer from the Cu/Cu^+ half-cell to the Zn/Zn^{2+} half-cell.
 - (D) Electrons will transfer from the Zn/Zn^{2+} half-cell to the Cu/Cu^+ half-cell.
11. For a reaction involving nitrogen monoxide inside a sealed flask, the value for the reaction quotient (Q) was found to be 1.1×10^2 at a given point. If, after this point, the amount of NO gas in the flask increased, which reaction is most likely taking place in the flask?
 - (A) $\text{NOBr}(g) \leftrightarrow \text{NO}(g) + \frac{1}{2}\text{Br}_2(g) \quad K_c = 3.4 \times 10^{-2}$
 - (B) $2\text{NOCl}(g) \leftrightarrow 2\text{NO}(g) + \text{Cl}_2(g) \quad K_c = 1.6 \times 10^{-5}$
 - (C) $2\text{NO}(g) + 2\text{H}_2(g) \leftrightarrow \text{N}_2(g) + 2\text{H}_2\text{O}(g) \quad K_c = 4.0 \times 10^6$
 - (D) $\text{N}_2(g) + \text{O}_2(g) \leftrightarrow 2\text{NO}(g) \quad K_c = 4.2 \times 10^2$
12. Which of the following substances has an asymmetrical molecular structure?
 - (A) SF_4
 - (B) PCl_5
 - (C) BF_3
 - (D) CO_2

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13.



The diagram above shows the speed distribution of molecules in a gas held at 200 K. Which of the following representations would best represent the gas at a higher temperature? (Note: The original line is shown as a dashed line in the answer options.)



14. Nitrogen's electronegativity value is between those of phosphorus and oxygen. Which of the following correctly describes the relationship between the three values?

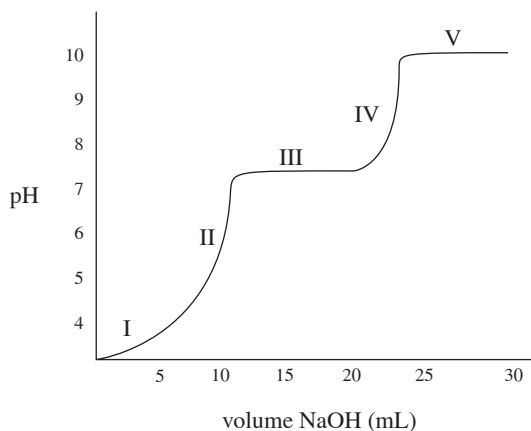
- (A) The value for nitrogen is less than that of phosphorus because nitrogen is larger, but greater than that of oxygen because nitrogen has a greater effective nuclear charge.
- (B) The value for nitrogen is less than that of phosphorus because nitrogen has fewer protons but greater than that of oxygen because nitrogen has fewer valence electrons.
- (C) The value for nitrogen is greater than that of phosphorus because nitrogen has fewer electrons, but less than that of oxygen because nitrogen is smaller.
- (D) The value for nitrogen is greater than that of phosphorus because nitrogen is smaller, but less than that of oxygen because nitrogen has a smaller effective nuclear charge.

15. A sample of a compound known to consist of only carbon, hydrogen, and oxygen is found to have a total mass of 29.05 g. If the mass of the carbon is 18.02 g and the mass of the hydrogen is 3.03 g, what is the empirical formula of the compound?

- (A) C_2H_4O
- (B) C_3H_6O
- (C) $C_2H_6O_3$
- (D) $C_3H_8O_2$

Questions 16-18 refer to the following.

A solution of carbonic acid, H_2CO_3 , is titrated with sodium hydroxide, NaOH. The following graph is produced:



16. In addition to OH^- , what species are present in the solution during section III of the graph?

- (A) H_2CO_3 , HCO_3^- , and CO_3^{2-}
- (B) H_2CO_3 and HCO_3^-
- (C) HCO_3^- and CO_3^{2-}
- (D) H_2CO_3 and CO_3^{2-}

17. What is the magnitude of the first dissociation constant?

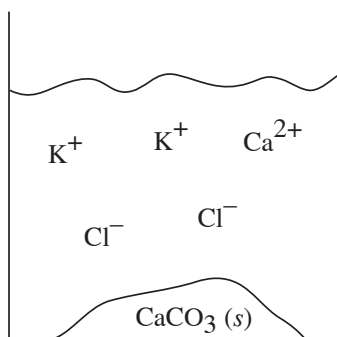
- (A) 10^{-2}
- (B) 10^{-4}
- (C) 10^{-6}
- (D) 10^{-8}

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18. If the concentration of the sodium hydroxide is increased prior to repeating the titration, what effect, if any, would that have on the graph?

(A) The graph would not change at all.
 (B) The pH values at the equivalence points would increase.
 (C) The equivalence points would be reached with less volume of NaOH added.
 (D) The slope of the equivalence points would decrease.

19. Solutions of potassium carbonate and calcium chloride are mixed, and the particulate representation below shows which are present in significant amounts after the reaction has gone to completion.



Which of the two original solutions is the limiting reagent and why?

- (A) The potassium carbonate, because of the polyatomic anion
 (B) The potassium carbonate, because there is no carbonate left after the reaction
 (C) The calcium chloride, because there is an excess of calcium ions post-reaction
 (D) The calcium chloride, because the component ions are smaller than those in potassium carbonate
20. In which of the following circumstances is the value for K_{eq} always greater than 1?

	ΔH	ΔS
(A)	Positive	Positive
(B)	Positive	Negative
(C)	Negative	Negative
(D)	Negative	Positive

21. The structure of two oxoacids is shown below:



Which would be a stronger acid, and why?

- (A) HOCl, because the H–O bond is weaker than in HOF as chlorine is larger than fluorine
 (B) HOCl, because the H–O bond is stronger than in HOF as chlorine has a higher electronegativity than fluorine
 (C) HOF, because the H–O bond is stronger than in HOCl as fluorine has a higher electronegativity than chlorine
 (D) HOF, because the H–O bond is weaker than in HOCl as fluorine is smaller than chlorine
22. During a chemical reaction, $\text{NO}(g)$ gets reduced and no nitrogen-containing compound is oxidized. Which of the following is a possible product of this reaction?
- (A) $\text{NO}_2(g)$
 (B) $\text{N}_2(g)$
 (C) $\text{NO}_3^-(aq)$
 (D) $\text{NO}_2^-(aq)$
23. Which of the following pairs of substances would make a good buffer solution?

- (A) $\text{HC}_2\text{H}_3\text{O}_2(aq)$ and $\text{NaC}_2\text{H}_3\text{O}_2(aq)$
 (B) $\text{H}_2\text{SO}_4(aq)$ and $\text{LiOH}(aq)$
 (C) $\text{HCl}(aq)$ and $\text{KCl}(aq)$
 (D) $\text{HF}(aq)$ and $\text{NH}_3(aq)$

Questions 24–27 refer to the following.

Inside a calorimeter, 100.0 mL of 1.0 *M* hydrocyanic acid (HCN), a weak acid, and 100.0 mL of 0.50 *M* sodium hydroxide are mixed. The temperature of the mixture rises from 21.5°C to 28.5°C. The specific heat of the mixture is approximately 4.2 J/g°C, and the density is identical to that of water.

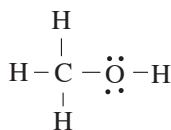
24. Identify the correct net ionic equation for the reaction that takes place.
- (A) $\text{HCN}(aq) + \text{OH}^-(aq) \leftrightarrow \text{CN}^-(aq) + \text{H}_2\text{O}(l)$
 (B) $\text{HCN}(aq) + \text{NaOH}(aq) \leftrightarrow \text{NaCN}(aq) + \text{H}_2\text{O}(l)$
 (C) $\text{H}^+(aq) + \text{OH}^-(aq) \leftrightarrow \text{H}_2\text{O}(l)$
 (D) $\text{H}^+(aq) + \text{CN}^-(aq) + \text{Na}^+(aq) + \text{OH}^-(aq) \leftrightarrow \text{H}_2\text{O}(l) + \text{CN}^-(aq) + \text{Na}^+(aq)$

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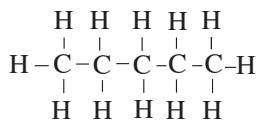
25. What is the approximate amount of heat released during the reaction?
- (A) 1.5 kJ
(B) 2.9 kJ
(C) 5.9 kJ
(D) 11.8 kJ
26. As ΔT increases, what happens to the equilibrium constant and why?
- (A) The equilibrium constant increases because more products are created.
(B) The equilibrium constant increases because the rate of the forward reaction increases.
(C) The equilibrium constant decreases because the equilibrium shifts to the left.
(D) The value for the equilibrium constant is unaffected by temperature and will not change.
27. If the experiment is repeated with 200.0 mL of 1.0 *M* HCN and 100. mL of 0.50 *M* NaOH, what would happen to the values for ΔT and ΔH_{rxn} ?

	ΔT	ΔH_{rxn}
(A)	Increase	Increase
(B)	Stay the same	Stay the same
(C)	Decrease	Stay the same
(D)	Stay the same	Increase

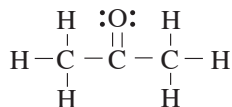
28. The following diagrams show the Lewis structures of four different molecules. Which molecule would travel the farthest in a paper chromatography experiment using a polar solvent?



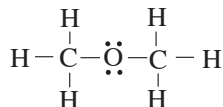
Methanol



Pentane



Acetone



Ether

- (A) Methanol
(B) Pentane
(C) Acetone
(D) Ether

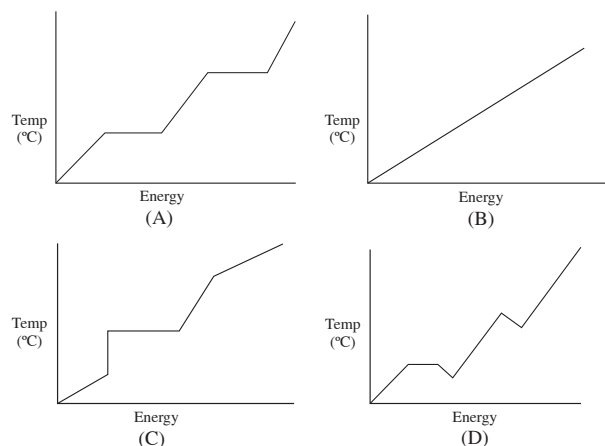
29. The first ionization energy for a neutral atom of chlorine is 1.25 MJ/mol and the first ionization energy for a neutral atom of argon is 1.52 MJ/mol. How would the first ionization energy value for a neutral atom of potassium compare to those values?
- (A) It would be greater than both because potassium carries a greater nuclear charge than either chlorine or argon.
(B) It would be greater than both because the size of a potassium atom is smaller than an atom of either chlorine or argon.
(C) It would be less than both because there are more electrons in potassium, meaning they repel each other more effectively and less energy is needed to remove one.
(D) It would be less than both because a valence electron of potassium is farther from the nucleus than one of either chlorine or argon.
30. Which net ionic equation below represents a possible reaction that takes place when a strip of magnesium metal is oxidized by a solution of chromium (III) nitrate?
- (A) $\text{Mg}(s) + \text{Cr}(\text{NO}_3)_3(aq) \rightarrow \text{Mg}^{2+}(aq) + \text{Cr}^{3+}(aq) + 3\text{NO}_3^-(aq)$
(B) $3\text{Mg}(s) + 2\text{Cr}^{3+} \rightarrow 3\text{Mg}^{2+} + 2\text{Cr}(s)$
(C) $\text{Mg}(s) + \text{Cr}^{3+} \rightarrow \text{Mg}^{2+} + \text{Cr}(s)$
(D) $3\text{Mg}(s) + 2\text{Cr}(\text{NO}_3)_3(aq) \rightarrow 3\text{Mg}^{2+}(aq) + 2\text{Cr}(s) + 6\text{NO}_3^-(aq)$
31. $\text{PCl}_3(g) + \text{Cl}_2(g) \leftrightarrow \text{PCl}_5(g) \quad \Delta H = -92.5 \text{ kJ/mol}$

In which of the following ways could the reaction above be manipulated to create more product?

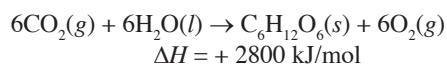
- (A) Decreasing the concentration of PCl_3
(B) Increasing the pressure
(C) Increasing the temperature
(D) None of the above

GO ON TO THE NEXT PAGE.

32. A pure solid substance is heated strongly. It first melts into a liquid, then boils and becomes a gas. Which of the following heating curves correctly shows the relationship between temperature and heat added?



33. Consider the following reaction showing photosynthesis:



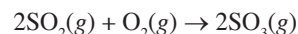
Which of the following is true regarding the thermal energy in this system?

- (A) It is transferred from the surroundings to the reaction.
 (B) It is transferred from the reaction to the surroundings.
 (C) It is transferred from the reactants to the products.
 (D) It is transferred from the products to the reactants.
34. $\text{SO}_2\text{Cl}_2 \rightarrow \text{SO}_2(g) + \text{Cl}_2(g)$

At 600 K, SO_2Cl_2 will decompose to form sulfur dioxide and chlorine gas via the above equation. If the reaction is found to be first order overall, which of the following will cause an increase in the half life of SO_2Cl_2 ?

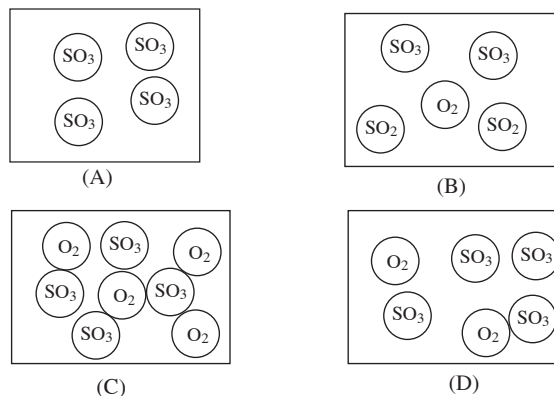
- (A) Increasing the initial concentration of SO_2Cl_2
 (B) Increasing the temperature at which the reaction occurs
 (C) Decreasing the overall pressure in the container
 (D) None of these will increase the half life.

Questions 35-38 refer to the following reaction.



4.0 mol of gaseous SO_2 and 6.0 mol of O_2 gas are allowed to react in a sealed container.

35. Which particulate drawing best represents the contents of the flask after the reaction goes to completion?

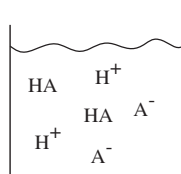


36. If the temperature remains constant, what percentage of the original pressure will the final pressure in the container be equal to?
- (A) 67%
 (B) 80%
 (C) 100%
 (D) 133%
37. At a given point in the reaction, all three gases are present at the same temperature. Which gas molecules will have the highest velocity and why?
- (A) The O_2 molecules, because they have the least mass
 (B) The O_2 molecules, because they are the smallest
 (C) The SO_3 molecules, because they are products in the reaction
 (D) Molecules of all three gases will have the same speed because they have the same temperature.

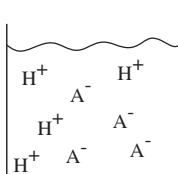
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38. Under which of the following conditions would the gases in the container most deviate from ideal conditions and why?
- (A) Low pressures, because the gas molecules would be spread far apart
- (B) High pressures, because the gas molecules will be colliding frequently
- (C) Low temperatures, because the intermolecular forces between the gas molecules would increase
- (D) High temperatures, because the gas molecules are moving too fast to interact with each other

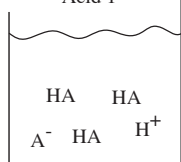
39. Four different acids are added to beakers of water, and the following diagrams represent the species present in each solution at equilibrium. Which acid has the highest pH?



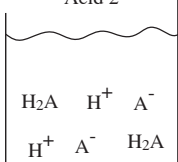
Acid 1



Acid 2



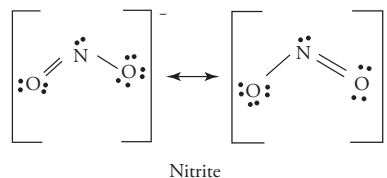
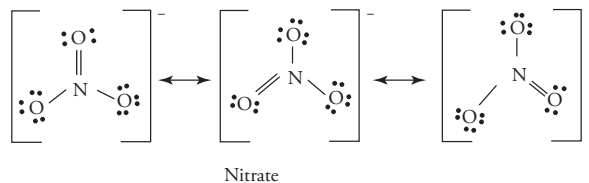
Acid 3



Acid 4

- (A) Acid 1
- (B) Acid 2
- (C) Acid 3
- (D) Acid 4
40. Which expression below should be used to calculate the mass of copper that can be plated out of a 1.0 M $\text{Cu}(\text{NO}_3)_2$ solution using a current of 0.75 A for 5.0 minutes?
- (A) $\frac{(5.0)(60)(0.75)(63.55)}{(96500)(2)}$
- (B) $\frac{(5.0)(60)(63.55)(2)}{(0.75)(96500)}$
- (C) $\frac{(5.0)(60)(96500)(0.75)}{(63.55)(2)}$
- (D) $\frac{(5.0)(60)(96500)(63.55)}{(0.75)(2)}$

41. Lewis diagrams for the nitrate and nitrite ions are shown below. Choose the statement that correctly describes the relationship between the two ions in terms of bond length and bond energy.



- (A) Nitrite has longer and stronger bonds than nitrate.
- (B) Nitrite has longer and weaker bonds than nitrate.
- (C) Nitrite has shorter and stronger bonds than nitrate.
- (D) Nitrite has shorter and weaker bonds than nitrate.
42. Examining data obtained from mass spectrometry supports which of the following?
- (A) The common oxidation states of elements
- (B) Atomic size trends within the periodic table
- (C) Ionization energy trends within the periodic table
- (D) The existence of isotopes
43. A 2.0 L flask holds 0.40 g of helium gas. If the helium is evacuated into a larger container while the temperature is held constant, what will the effect on the entropy of the helium be?
- (A) It will remain constant as the number of helium molecules does not change.
- (B) It will decrease as the gas will be more ordered in the larger flask.
- (C) It will decrease because the molecules will collide with the sides of the larger flask less often than they did in the smaller flask.
- (D) It will increase as the gas molecules will be more dispersed in the larger flask.

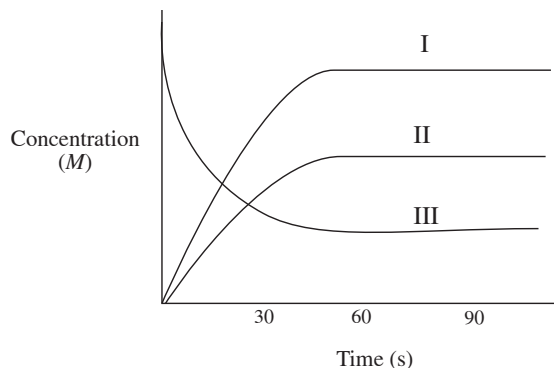
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Questions 44–48 refer to the following.

NO_2 gas is placed in a sealed, evacuated container and allowed to decompose via the following equation:



The graph below indicates the change in concentration for each species over time.



44. Using the numbers of the lines on the graph, identify which line belongs to which species.

	Line I	Line II	Line III
(A)	NO_2	NO	O_2
(B)	O_2	NO_2	NO
(C)	NO	O_2	NO_2
(D)	O_2	NO	NO_2

45. What is happening to the rate of the forward reaction at $t = 60$ s?

- (A) It is increasing.
 (B) It is decreasing.
 (C) It is remaining constant.
 (D) It is zero.

46. As the reaction progresses, what happens to the value of the equilibrium constant K_p if the temperature remains constant?

- (A) It stays constant.
 (B) It increases exponentially.
 (C) It increases linearly.
 (D) It decreases exponentially.

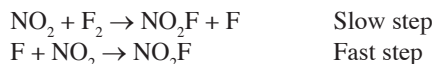
47. What would happen to the slope of the NO_2 line if additional O_2 were injected into the container?

- (A) It would increase, then level off.
 (B) It would decrease, then level off.
 (C) It would remain constant.
 (D) It would increase, then decrease.

48. Using the graph, how could you determine the instantaneous rate of disappearance of NO_2 at $t = 30$ s?

- (A) By determining the area under the graph at $t = 30$ s
 (B) By taking the slope of a line tangent to the NO_2 curve at $t = 30$ s
 (C) By using the values at $t = 30$ s and plugging them into the K_p expression
 (D) By measuring the overall gas pressure in the container at $t = 30$ s

49. A proposed mechanism for a reaction is as follows:



What is the order of the overall reaction?

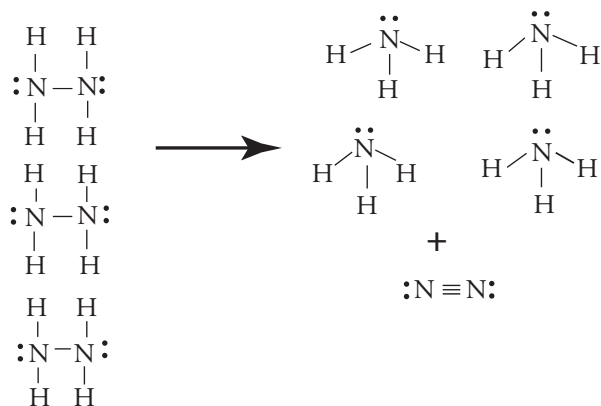
- (A) Zero order
 (B) First order
 (C) Second order
 (D) Third order

50. Starting with a stock solution of $18.0\text{ M H}_2\text{SO}_4$, what is the proper procedure to create a 1.00 L sample of a 3.0 M solution of H_2SO_4 in a volumetric flask?

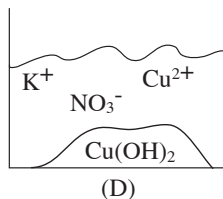
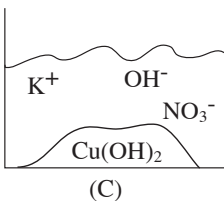
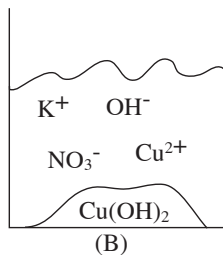
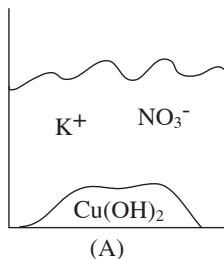
- (A) Add 167 mL of the stock solution to the flask, then fill the flask the rest of the way with distilled water while swirling the solution.
 (B) Add 600 mL of the stock solution to the flask, then fill the flask the rest of the way with distilled water while swirling the solution.
 (C) Fill the flask partway with water, then add 167 mL of the stock solution, swirling to mix it. Last, fill the flask the rest of the way with distilled water.
 (D) Fill the flask partway with water, then add 600 mL of the stock solution, swirling to mix it. Last, fill the flask the rest of the way with distilled water.

GO ON TO THE NEXT PAGE.

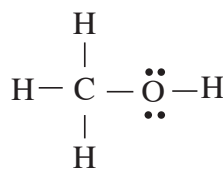
51. The reaction shown in the diagram below is accompanied by a large increase in temperature. If all molecules shown are in their gaseous state, which statement accurately describes the reaction?



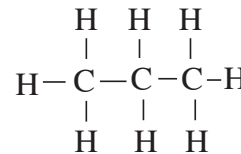
- (A) It is an exothermic reaction in which entropy increases.
 (B) It is an exothermic reaction in which entropy decreases.
 (C) It is an endothermic reaction in which entropy increases.
 (D) It is an endothermic reaction in which entropy decreases.
52. A student mixes equimolar amounts of KOH and $\text{Cu}(\text{NO}_3)_2$ in a beaker. Which of the following particulate diagrams correctly shows all species present after the reaction occurs?



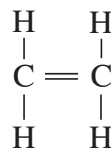
Questions 53-55 refer to the following.



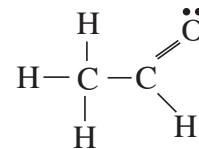
Methanol



Propane



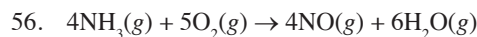
Ethene



Ethanal

53. Based on the strength of the intermolecular forces in each substance, estimate from greatest to smallest the vapor pressures of each substance in liquid state at the same temperature.
- (A) Propane > Ethanal > Ethene > Methanol
 (B) Ethene > Propane > Ethanal > Methanol
 (C) Ethanal > Methanol > Ethene > Propane
 (D) Methanol > Ethanal > Propane > Ethene
54. When in liquid state, which two substances are most likely to be miscible with water?
- (A) Propane and ethene
 (B) Methanol and propane
 (C) Ethene and ethanal
 (D) Methanol and ethanal
55. Between propane and ethene, which will likely have the higher boiling point and why?
- (A) Propane, because it has a greater molar mass
 (B) Propane, because it has a more polarizable electron cloud
 (C) Ethene, because of the double bond
 (D) Ethene, because it is smaller in size

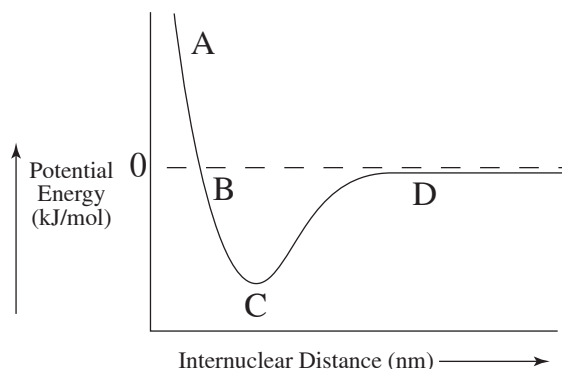
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The above reaction will experience a rate increase by the addition of a catalyst such as platinum. Which of the following best explains why?

- (A) The catalyst causes the value for ΔG to become more negative.
- (B) The catalyst increases the percentage of collisions that occur at the proper orientation in the reactant molecules.
- (C) The catalyst introduces a new reaction mechanism for the reaction.
- (D) The catalyst increases the activation energy for the reaction.

57. The graph below shows the amount of potential energy between two hydrogen atoms as the distance between them changes. At which point in the graph would a molecule of H_2 be the most stable?



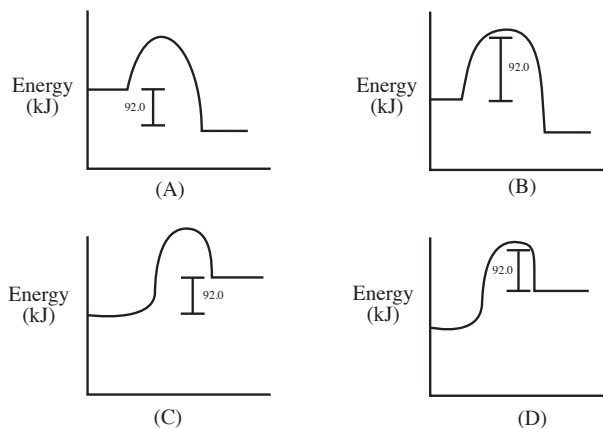
- (A) Point A
 - (B) Point B
 - (C) Point C
 - (D) Point D
58. $\text{N}_2(g) + \text{O}_2(g) + \text{Cl}_2(g) \leftrightarrow 2\text{NOCl}(g) \quad \Delta G^\circ = 132.6 \text{ kJ/mol}$

For the equilibrium above, what would happen to the value of ΔG° if the concentration of N_2 were to increase and why?

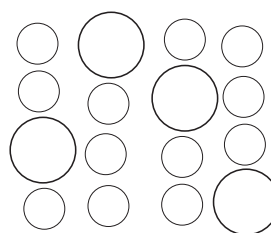
- (A) It would increase as the reaction would become more thermodynamically favored.
- (B) It would increase as the reaction would shift right and create more products.
- (C) It would decrease because there are more reactants present.
- (D) It would stay the same because the value of K_{eq} would not change.



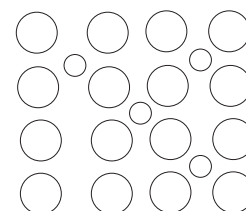
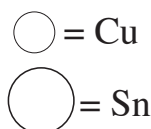
Which of the following energy level diagrams gives an accurate representation of the above reaction?



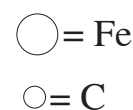
60. Two alloys are shown in the diagrams below—bronze, and steel. Which of the following correctly describes the malleability of both alloys compared to their primary metals?



Bronze



Steel



- (A) Bronze's malleability would be comparable to that of copper, but steel's malleability would be significantly lower than that of iron.
- (B) Bronze's malleability would be significantly higher than that of copper, but steel's malleability would be comparable to that of iron.
- (C) Both bronze and steel would have malleability values similar to those of their primary metals.
- (D) Both bronze and steel would have malleability values lower than those of their primary metals.

END OF SECTION I

CHEMISTRY**SECTION II****Time—1 hour and 45 minutes**General Instructions

Calculators, including those with programming and graphing capabilities, may be used. However, calculators with typewriter-style (QWERTY) keyboards are NOT permitted.

Pages containing a periodic table and equations commonly used in chemistry will be available for your use.

You may write your answers with either a pen or a pencil. Be sure to write CLEARLY and LEGIBLY. If you make an error, you may save time by crossing it out rather than trying to erase it.

Write all your answers in the essay booklet. Number your answers as the questions are numbered in the examination booklet.

GO ON TO THE NEXT PAGE.

INFORMATION IN THE TABLE BELOW AND ON THE FOLLOWING PAGES MAY BE USEFUL IN ANSWERING THE QUESTIONS IN THIS SECTION OF THE EXAMINATION.

DO NOT DETACH FROM BOOK.

PERIODIC TABLE OF THE ELEMENTS

PERIODIC TABLE OF THE ELEMENTS																							
1	2																						
H	He																						
1.008	4.00																						
3	4															9	10						
Li	Be															F	Ne						
6.94	9.01															19.00	20.18						
11	12															17	18						
Na	Mg															Cl	Ar						
22.99	24.30															35.45	39.95						
19	20	21	22	23	24	25	26	27	28	29	30					31	32	33	34	35	36		
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn					Ga	Ge	As	Se	Br	Kr		
39.10	40.08	44.96	47.90	50.94	52.00	54.94	55.85	58.93	58.69	63.55	65.39					69.72	72.59	74.92	78.96	79.90	83.80		
37	38	39	40	41	42	43	44	45	46	47	48					49	50	51	52	53	54		
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd					In	Sn	Sb	Te	I	Xe		
85.47	87.62	88.91	91.22	92.91	95.94	(98)	101.1	102.91	106.42	107.87	112.41					114.82	118.71	121.75	127.60	126.91	131.29		
55	56	57	72	73	74	75	76	77	78	79	80					81	82	83	84	85	86		
Cs	Ba	*La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg					Tl	Pb	Bi	Po	At	Rn		
132.91	137.33	138.91	178.49	180.95	183.85	186.21	190.2	192.2	195.08	196.97	200.59					204.38	207.2	208.98	(209)	(210)	(222)		
87	88	89	104	105	106	107	108	109	110	111													
Fr	Ra	†Ac	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg													
(223)	226.02	227.03	(261)	(262)	(266)	(264)	(277)	(268)	(271)	(272)													

ADVANCED PLACEMENT CHEMISTRY EQUATIONS AND CONSTANTS

Throughout the test the following symbols have the definitions specified unless otherwise noted.

L, mL = liter(s), milliliter(s)
 g = gram(s)
 nm = nanometer(s)
 atm = atmosphere(s)

mm Hg = millimeters of mercury
 J, kJ = joule(s), kilojoule(s)
 V = volt(s)
 mol = mole(s)

ATOMIC STRUCTURE

$$E = h\nu$$

$$c = \lambda\nu$$

E = energy
 ν = frequency
 λ = wavelength

Planck's constant, $h = 6.626 \times 10^{-34} \text{ J s}$

Speed of light, $c = 2.998 \times 10^8 \text{ m s}^{-1}$

Avogadro's number = $6.022 \times 10^{23} \text{ mol}^{-1}$

Electron charge, $e = -1.602 \times 10^{-19} \text{ coulombs}$

EQUILIBRIUM

$$K_c = \frac{[C]^c[D]^d}{[A]^a[B]^b}, \text{ where } aA + bB \rightleftharpoons cC + dD$$

$$K_p = \frac{(P_C)^c(P_D)^d}{(P_A)^a(P_B)^b}$$

$$K_a = \frac{[H^+][A^-]}{[HA]}$$

$$K_b = \frac{[OH^-][HB^+]}{[B]}$$

$$K_w = [H^+][OH^-] = 1.0 \times 10^{-14} \text{ at } 25^\circ\text{C}$$

$$= K_a \times K_b$$

$$\text{pH} = -\log[H^+], \text{ pOH} = -\log[OH^-]$$

$$14 = \text{pH} + \text{pOH}$$

$$\text{pH} = \text{p}K_a + \log \frac{[A^-]}{[HA]}$$

$$\text{p}K_a = -\log K_a, \text{ p}K_b = -\log K_b$$

Equilibrium Constants

K_c (molar concentrations)

K_p (gas pressures)

K_a (weak acid)

K_b (weak base)

K_w (water)

KINETICS

$$\ln[A]_t - \ln[A]_0 = -kt$$

$$\frac{1}{[A]_t} - \frac{1}{[A]_0} = kt$$

$$t_{1/2} = \frac{0.693}{k}$$

k = rate constant

t = time

$t_{1/2}$ = half-life

GO ON TO THE NEXT PAGE.

GASES, LIQUIDS, AND SOLUTIONS

$$PV = nRT$$

$$P_A = P_{\text{total}} \times X_A, \text{ where } X_A = \frac{\text{moles A}}{\text{total moles}}$$

$$P_{\text{total}} = P_A + P_B + P_C + \dots$$

$$n = \frac{m}{M}$$

$$K = ^\circ\text{C} + 273$$

$$D = \frac{m}{V}$$

$$KE \text{ per molecule} = \frac{1}{2}mv^2$$

Molarity, M = moles of solute per liter of solution

$$A = abc$$

P = pressure

V = volume

T = temperature

n = number of moles

m = mass

M = molar mass

D = density

KE = kinetic energy

v = velocity

A = absorbance

a = molar absorptivity

b = path length

c = concentration

$$\begin{aligned} \text{Gas constant, } R &= 8.314 \text{ J mol}^{-1} \text{ K}^{-1} \\ &= 0.08206 \text{ L atm mol}^{-1} \text{ K}^{-1} \\ &= 62.36 \text{ L torr mol}^{-1} \text{ K}^{-1} \end{aligned}$$

$$1 \text{ atm} = 760 \text{ mm Hg}$$

$$= 760 \text{ torr}$$

$$\text{STP} = 0.00^\circ\text{C and } 1.000 \text{ atm}$$

THERMOCHEMISTRY/ ELECTROCHEMISTRY

$$q = mc\Delta T$$

$$\Delta S^\circ = \sum S^\circ \text{ products} - \sum S^\circ \text{ reactants}$$

$$\Delta H^\circ = \sum \Delta H_f^\circ \text{ products} - \sum \Delta H_f^\circ \text{ reactants}$$

$$\Delta G^\circ = \sum \Delta G_f^\circ \text{ products} - \sum \Delta G_f^\circ \text{ reactants}$$

$$\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$$

$$= -RT \ln K$$

$$= -nFE^\circ$$

$$I = \frac{q}{t}$$

q = heat

m = mass

c = specific heat capacity

T = temperature

S° = standard entropy

H° = standard enthalpy

G° = standard free energy

n = number of moles

E° = standard reduction potential

I = current (amperes)

q = charge (coulombs)

t = time (seconds)

Faraday's constant, F = 96,485 coulombs per mole of electrons

$$1 \text{ volt} = \frac{1 \text{ joule}}{1 \text{ coulomb}}$$

GO ON TO THE NEXT PAGE.

CHEMISTRY

Section II

(Total time—105 minutes)

YOU MAY USE YOUR CALCULATOR IN THIS SECTION

THE METHODS USED AND THE STEPS INVOLVED IN ARRIVING AT YOUR ANSWERS MUST BE SHOWN CLEARLY. It is to your advantage to do this since you may obtain partial credit if you do, and you will receive little or no credit if you do not. Attention should be paid to significant figures.

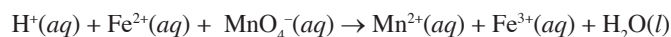
1. A student is tasked with determining the identity of an unknown carbonate compound with a mass of 1.89 g. The compound is first placed in water, where it dissolves completely. The K_{sp} value for several carbonate-containing compounds are given below.

Compound	K_{sp}
Lithium carbonate	8.15×10^{-4}
Nickel (II) carbonate	1.42×10^{-7}
Strontium carbonate	5.60×10^{-10}

- In order to precipitate the maximum amount of the carbonate ions from solution, which of the following should be added to the carbonate solution: LiNO_3 , $\text{Ni}(\text{NO}_3)_2$, or $\text{Sr}(\text{NO}_3)_2$? Justify your answer.
- For the carbonate compound that contains the cation chosen in part (a), determine the concentration of each ion of that compound in solution at equilibrium.
- When mixing the solution, should the student ensure the carbonate solution or the nitrate solution is in excess? Justify your answer.
- After titrating sufficient solution to precipitate out all of the carbonate ions, the student filters the solution before placing it in a crucible and heating it to drive off the water. After several heatings, the final mass of the precipitate remains constant and is determined to be 2.02 g.
 - Determine the number of moles of precipitate.
 - Determine the mass of carbonate present in the precipitate.
- Determine the percent, by mass, of carbonate in the original sample.
- Is the original compound most likely lithium carbonate, sodium carbonate, or potassium carbonate? Justify your answer.

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2. The **unbalanced** reaction between potassium permanganate and acidified iron (II) sulfate is a redox reaction that proceeds as follows:



- (a) Provide the equations for both half-reactions that occur below:
 (i) Oxidation half-reaction
 (ii) Reduction half-reaction
 (b) What is the balanced net ionic equation?

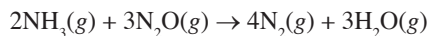
A solution of 0.150 *M* potassium permanganate is placed in a buret before being titrated into a flask containing 50.00 mL of iron (II) sulfate solution of unknown concentration. The following data describes the colors of the various ions in solution:

Ion	Color in solution
H^+	Colorless
Fe^{2+}	Pale Green
MnO_4^-	Dark Purple
Mn^{2+}	Colorless
Fe^{3+}	Yellow
K^+	Colorless
SO_4^{2-}	Colorless

- (c) Describe the color of the solution in the flask at the following points:
 (i) Before titration begins
 (ii) During titration prior to the endpoint
 (iii) At the endpoint of the titration
 (d) (i) If 15.55 mL of permanganate are added to reach the endpoint, what is the initial concentration of the iron (II) sulfate?
 (ii) The actual concentration of the FeSO_4 is 0.250 *M*. Calculate the percent error.
 (e) Could the following errors have led to the experimental result deviating in the direction that it did? You must justify your answers quantitatively.
 (i) 55.0 mL of FeSO_4 was added to the flask prior to titration instead of 50.0 mL.
 (ii) The concentration of the potassium permanganate was actually 0.160 *M* instead of 0.150 *M*.

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3. Ammonia gas reacts with dinitrogen monoxide via the following reaction:



The absolute entropy values for the varying substances are listed in the table below.

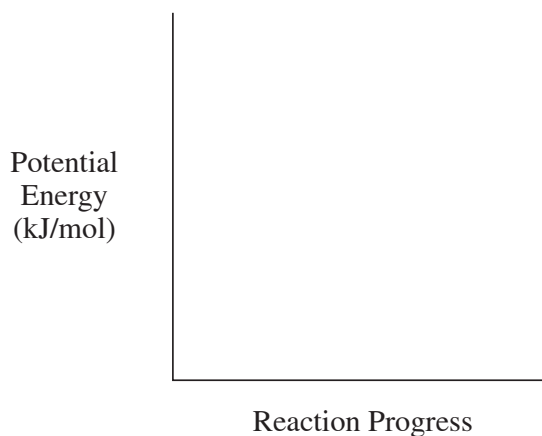
Substance	S° (J/mol·K)
$\text{NH}_3(g)$	193
$\text{N}_2\text{O}(g)$	220
$\text{N}_2(g)$	192
$\text{H}_2\text{O}(g)$	189

- (a) Calculate the entropy value for the overall reaction.

Several bond enthalpies are listed in the table below.

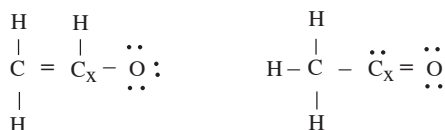
Bond	Enthalpy (kJ/mol)	Bond	Enthalpy (kJ/mol)
N–H	388	N=N	409
N–O	210	N≡N	941
N=O	630	O–H	463

- (b) Calculate the enthalpy value for the overall reaction.
 (c) Is this reaction thermodynamically favored at 25°C? Justify your answer.
 (d) If 25.00 g of NH_3 reacts with 25.00 g of N_2O :
 (i) Will energy be released or absorbed?
 (ii) What is the magnitude of the energy change?
 (e) On the reaction coordinates below, draw a line showing the progression of this reaction. Label both ΔH and E_a on the graph.



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4. The acetyl ion has a formula of $C_2H_3O^-$ and two possible Lewis electron-dot diagram representations:



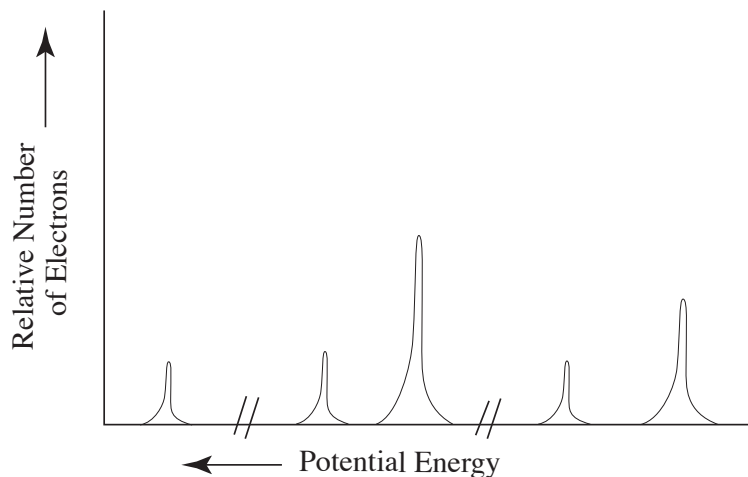
- (a) Using formal charge, determine which structure is the most likely correct structure.
- (b) For carbon atom “x” in the structure you chose:
- What is the hybridization around the atom?
 - How many sigma and pi bonds has the atom formed?
- (c) A hydrogen ion attaches itself to the the acetyl ion, creating C_2H_4O . Draw the Lewis diagram of the new molecule.
5. A student tests the conductivity of three different acid samples, each with a concentration of 0.10 *M* and a volume of 20.0 mL. The conductivity was recorded in microsiemens per centimeter in the table below:

Sample	Conductivity ($\mu\text{S}/\text{cm}$)
1	26,820
2	8655
3	35,120

- (a) The three acids are known to be HCl , H_2SO_4 , and H_3PO_4 . Identify which sample is which acid. Justify your answer.
- (b) The HCl solution is then titrated with a 0.150 *M* solution of the weak base methylamine, CH_3NH_2 . ($K_b = 4.38 \times 10^{-4}$)
- Write out the net ionic equation for this reaction.
 - Determine the pH of the solution after 20.0 mL of methylamine has been added.

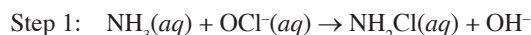
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6. The photoelectron spectrum of an element is given below:



- (a) Identify the element this spectra most likely belongs to and write out its full electron configuration.
- (b) Using your knowledge of atomic structure, explain the following:
 - (i) The reason for the three discrete areas of ionization energies
 - (ii) The justification for there being a total of five peaks
 - (iii) The relative heights of the peaks when compared to each other

7. Hydrazine (N_2H_4) can be produced commercially via the Raschig process. The following is a proposed mechanism:



- (a)
 - (i) What is the equation for the overall reaction?
 - (ii) Identify any catalysts or intermediates from the reaction mechanism.
- (b) The rate law for the reaction is determined to be $\text{rate} = k[\text{NH}_3][\text{OCl}^-]$
 - (i) Which elementary step is the slowest one? Justify your answer.
 - (ii) If the reaction is measured over the course of several minutes, what would the units of the rate constant be?

STOP

END OF EXAM

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(Print) Last First M.I.

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(Print) Number and Street

City State Zip Code

PHONE NO.: _____

IMPORTANT: Please fill in these boxes exactly as shown on the back cover of your test book.

2. TEST FORM

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4	E	N	4	4	4	4	4	4	4	4
5	F	O	5	5	5	5	5	5	5	5
6	G	P	6	6	6	6	6	6	6	6
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4. REGISTRATION NUMBER

7. GENDER

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5. YOUR NAME

First 4 letters of last name				FIRST INIT	MID INIT
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G	G	G	G	G	G
H	H	H	H	H	H
I	I	I	I	I	I
J	J	J	J	J	J
K	K	K	K	K	K
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M	M	M	M	M	M
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